

Molecular Shapes Assignment

DROPBOX: Molecular Shapes Assignment

1. Predict the Lewis structure and bond angles for each of the following molecules:
 - a. $C_2Br_4(s)$
 - b. $HgCl_2(s)$
 - c. $SnCl_4(aq)$
2. Using VSEPR Theory, sketch the molecular shape of each of the molecules listed in question #1.
3. Draw the Lewis Structure of a sulfur dioxide molecule and use the VSEPR theory to predict its shape and bond angle.
4. Use the hybridization theory to account for the bonding and shape of an ammonia molecule.
5. Explain the bonding and shape of ethyne, $C_2H_2(g)$, using the hybridization theory.

GO TO

dropbox to view the rubric and upload your assignment in order to receive feedback from your teacher.

OUT OF
100

Assessment OF: This assignment will be evaluated for a grade or mark that will contribute to your overall final mark in this course.

